Paper Code: _____

1	1
2	2
3	3

ESKP-09xvii01

CHEMISTRY 9TH

P-198

Time: 3 hours

Marks: 65

Note: There are THREE sections in this paper i.e. A, B & C. Attempt Section-A and return it to the Superintendent within the given time. No marks will be awarded for cutting, erasing and overwriting.

Time: 15 minutes **Section-A** **Marks: 12**

QNo.1 Select the correct option and insert (A,B,C,D) in the relevant box.

i. One atomic mass unit (amu) is _____. B

A- 1.67×10^{-25}g B- 1.67×10^{-24}g C- 1.67×10^{-23}g D- 1.67×10^{-22}g

ii. Which one of the following is a homogenous mixture? B

A- Smoke B- Air C- Fog D- Smog

iii. Carbon has _____ isotopes. B

A- Five B- Three C- Two D- Four

iv. Alpha particles are _____. D

A- Neutral B- Negatively charged C- Protons D- Double positively charged

v. The group I and group II elements are called the _____ elements. B

A- P-block B- S-block C- D-block D- F-block

vi. The modern periodic table is based on the _____. D

A- Law of triads B- Law of octaves C- Mendeleev's Periodic Law D- Modern Periodic Law

vii. Which kind of bond exists in HCl? C

A- Covalent bond B- Ionic bond C- Polar bond D- Coordinate bond

viii. When the temperature of water is dropped to _____, it changes into ice. C

A- 100°C B- 50°C C- 0°C D- 4°C

ix. The increase in temperature of the gases decreases the _____. C

A- Pressure B- Volume C- Force of attraction D- Kinetic energy

x. Solution whose concentration is known is called _____. C

A- Aqueous solution B- Saturated solution C- Standard solution D- Non-standard solution

xi. Which one of the following is a strong electrolyte in solution? C

A- H_2CO_3 B- NH_4OH C- HCl D- CH_3COOH

xii. The halogens are _____. B

A- Inert B- Reactive C- Electropositive D- Metallic

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Note: Time allowed 2:45 hours.

SECTION – B

Marks: 32

Q2: Answer any EIGHT parts. Each part caries equal marks.

i. Differentiate between Analytical and Bio Chemistry.

ii. Calculate the molecular mass of glucose ($C_6H_{12}O_6$), the atomic mass of C, H and O to be 12, 1, 16 respectively.

iii. Write two de-merits of Ratherford Atomic Model.

iv. Why atom is a neutral particle?

v. Define Sheilding effect. Does it vary in a period?

vi. What is metallic bond?

vii. Differentiate between Crystalline solids and Amorphous solid.

viii. It takes long time to cook at high altitude. Why?

ix. Define saturated and unsaturated solutions.

x. Define Oxidation and Reduction.

xi. What is aqua regia?

SECTION – C

Marks: 21

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Note: Attempt any THREE of the following. All questions carry equal marks.

Q3. (a) How many grams are there in 3.5 moles of water?

(b) Give the electronic configuration of elements having atomic number 5 to 10.

Q4. (a) Define the electronegativity of an element. Discuss its periodic variation in a period and in a group in the periodic table.

(b) Define covalent bond and explain its types in detail.

Q5. (a) What is vapour pressure? What are the factors, which effect vapour pressure?

(b) Define and explain the following with examples:
(i) Suspension (ii) Colloidal solution

Q6. (a) Describe Daniel cell.

(b) Write down chemical properties of Halogens.