

Roll No

Name

1- ہر سوال کے سامنے چار دائرے دئے گئے ہیں، صرف صحیح جواب والا دائرہ بھریں۔

2- دائروں کو شیڈ (بھرنے) کے لئے نیلے یا کالے رنگ کا پین استعمال کریں۔

3- جواب میں ایک سے زائد دائرے بھرنے سے جواب غلط تصور ہوگا۔



Time Allowed: 20 Minutes

SECTION - A

Marks : 18

Kp is more than Kc when the difference of the moles of the products and reactants is	<input type="radio"/> Negative	<input type="radio"/> One	<input type="radio"/> Zero	<input checked="" type="radio"/> Positive
Depression of freezing point for one Molal solution of a non volatile non electrolyte is called constant.	<input type="radio"/> Ebullioscopic	<input type="radio"/> Rate	<input type="radio"/> Equilibrium	<input checked="" type="radio"/> None of these
A cubic crystal has..... centre/s of symmetry.	<input type="radio"/> Three	<input type="radio"/> Six	<input type="radio"/> Two	<input checked="" type="radio"/> One
About 99 % of the universe is made up of	<input type="radio"/> Solid	<input checked="" type="radio"/> Plasma	<input type="radio"/> Liquid	<input type="radio"/> Gasses
Paschan series is obtained near region.	<input type="radio"/> U.V	<input checked="" type="radio"/> IR	<input type="radio"/> Visible	<input type="radio"/> Microwave
The oxidation number of Manganese in MnO_4^{-1} is.....	<input checked="" type="radio"/> +7	<input type="radio"/> -1	<input type="radio"/> +3	<input type="radio"/> +5
The total no of orbital in the second orbit is	<input type="radio"/> 2	<input type="radio"/> 4	<input type="radio"/> 6	<input checked="" type="radio"/> 8
The no of molecules in 125 gm of hydrated copper sulphate is	<input type="radio"/> 3.011×10^{23}	<input type="radio"/> 6.02×10^{23}	<input type="radio"/> 10.2×10^{23}	<input checked="" type="radio"/> None of these
At the same temperature and pressure, which of the given gases has the greatest density?	<input checked="" type="radio"/> Cl_2	<input type="radio"/> H_2O	<input type="radio"/> CO_2	<input type="radio"/> SO_2
For a reaction $A \longrightarrow$ product, doubling the concentration of A, quadruples the rate. The reaction is ...	<input type="radio"/> First order	<input checked="" type="radio"/> 2 nd order	<input type="radio"/> 3 rd order	<input type="radio"/> Zero order
Enthalpy means.....	<input checked="" type="radio"/> Heat constant	<input type="radio"/> Rate constant	<input type="radio"/> Disorder	<input type="radio"/> Transition state
The amount of heat required to break one mole of H_2 bond into individual atom is	<input type="radio"/> 425 KJ	<input type="radio"/> 430 KJ	<input type="radio"/> 436 KJ	<input type="radio"/> 440 KJ
The percentage composition of carbon in glucose is	<input type="radio"/> 6.67 %	<input checked="" type="radio"/> 40 %	<input type="radio"/> 45.5 %	<input type="radio"/> None of these
Born Haber cycle is a special application of law.	<input type="radio"/> Boyles	<input type="radio"/> Rate	<input checked="" type="radio"/> Hess's	<input type="radio"/> Hennery
When frequency of wave is higher then its energy will be.....	<input type="radio"/> Higher	<input type="radio"/> Lower	<input type="radio"/> Unknown	<input type="radio"/> Independent
Which of the given is not electromagnetic in nature?	<input checked="" type="radio"/> γ - rays	<input type="radio"/> X - rays	<input type="radio"/> Cathode rays	<input type="radio"/> IR rays
Activated complex is a substance which	<input type="radio"/> Is Stable	<input checked="" type="radio"/> Is Unstable	<input type="radio"/> Can be isolated	<input type="radio"/> Is Normal
SI unit of coefficient of viscosity is.....	<input type="radio"/> Kg/m	<input type="radio"/> Kg/s	<input type="radio"/> Kg/m^2	<input checked="" type="radio"/> Kg/ms

PR XI (01) 17 1-200
CHEMISTRY (New)
Inter Part-I
(Fresh/Reappear)

Note: Time allowed for Section – B and Section – C is 2 Hours and 40 minutes.

Section – B

Marks: 40

Q-II Answer any TEN parts. Each part carries FOUR marks.

1. Explain briefly the concept of mole.
2. How will you identify the limiting reagent in a chemical reaction?
3. Describe any four properties of cathode rays.
4. Describe briefly the Roentgen Method for the production of X – rays.
5. CO₂ molecule has linear structure, but H₂O molecule has angular structure. Comment.
6. Do you think that some postulates of Kinetic Molecular Theory are fault, write down these postulate.
7. Write down any four applications of Hydrogen Bonding.
8. Define and differentiate between crystalline and amorphous solids.
9. Discuss the Equilibrium of a sparingly soluble salt.
10. Prove that pH + pOH = 14.
11. Write down the relationship between reactant concentration and rate of reaction.
12. Write rules for assigning the oxidation number.
13. Write a note on fuel cells.

Section – C

Marks: 27

Note : Attempt any THREE questions. All questions carry equal marks.

- Q-III (A) Write down the postulates of Planck's Quantum Theory. 04
(B) Differentiate between continuous and line spectrum. 05
- Q-IV (A) Discuss the reactions of ionic and covalent compounds. 04
(B) State Graham's Law of diffusion. How can it be experimentally verified. 05
- Q-V (A) What do you know about dipole-dipole interactions? 04
(B) Discuss briefly the Collision Theory of reaction rate. 05
- Q-VI (A) What are Buffer solutions? Discuss briefly. 04
(B) Balance the following Equation by oxidation Number method. 05

